

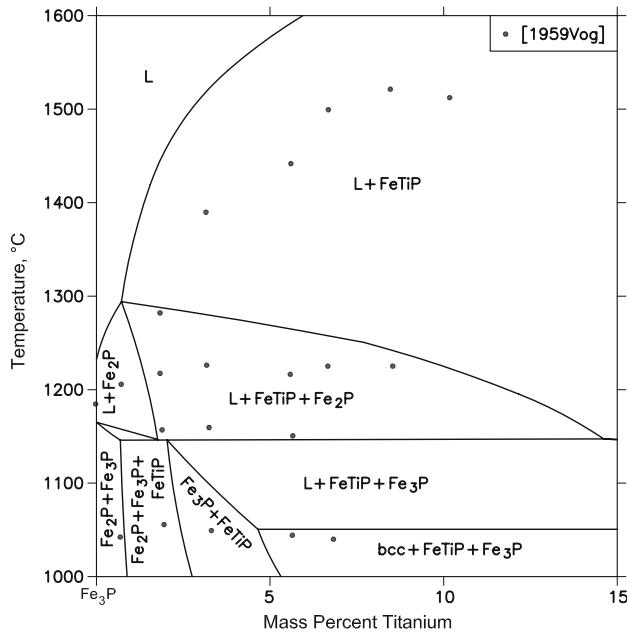
# Fe-P-Ti (Iron-Phosphorus-Titanium)

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The previous review of the experimental data on this ternary system by [1988Rag], based on mainly on the work of [1959Vog] and [1965Kan], presented a liquidus projection, an isothermal section at 25 °C and a reaction scheme for the Fe-FeP-Ti region. Recently, [2006Oht] made a CALPHAD-type thermodynamic analysis of the entire system, incorporating first-principles energy calculations.

## Binary Systems

The Fe-P phase diagram [2006Oht, Massalski2] has the following compounds at ambient pressure:  $\text{Fe}_3\text{P}$  ( $D_{0e}$ ,  $\text{Ni}_3\text{P}$ -type tetragonal),  $\text{Fe}_2\text{P}$  ( $C22$ ,  $\text{Fe}_2\text{P}$ -type hexagonal),  $\text{FeP}$  ( $B31$ ,  $\text{MnP}$ -type orthorhombic),  $\text{FeP}_2$  ( $C18$ ,  $\text{FeS}_2$ -marcasite type orthorhombic), and  $\text{FeP}_4$  (monoclinic). There are two intermediate phases in Fe-Ti system:  $\text{Fe}_2\text{Ti}$  ( $C14$ ,  $\text{MgZn}_2$ -type hexagonal) and  $\text{FeTi}$  ( $B2$ ,  $\text{CsCl}$ -type cubic). The computed P-Ti phase diagram [2006Oht] shows the following intermediate phases:  $\text{Ti}_3\text{P}$  (tetragonal, space group  $P4_2/n$ ),  $\text{Ti}_2\text{P}$  (hexagonal, space group  $P6/mmm$ ),  $\text{Ti}_{1.7}\text{P}$  (orthorhombic, space group  $P2_12_12_1$ ),  $\text{Ti}_5\text{P}_3$  ( $D8_8$ ,  $\text{Mn}_5\text{Si}_3$ -type hexagonal),  $\text{Ti}_4\text{P}_3$  ( $D7_3$ ,  $\text{Th}_3\text{P}_4$ -type cubic?),  $\text{TiP}$  ( $B_i$ ,  $\text{AsTi}$ -type hexagonal), and  $\text{TiP}_2$  (orthorhombic, space group  $Pnma$ ).

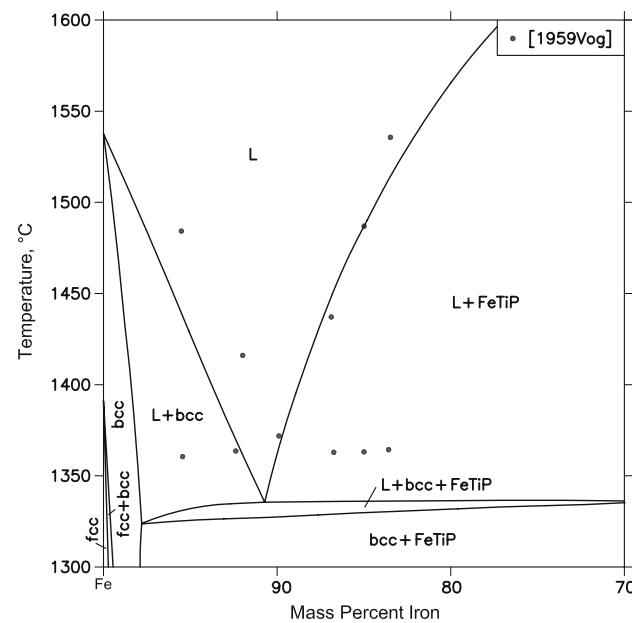


**Fig. 1** Fe-P-Ti computed vertical section along the  $\text{Fe}_3\text{P}$ - $\text{FeTiP}$  join [2006Oht]

## The Computed Ternary Phase Equilibria

In their thermodynamic modeling, [2006Oht] used a regular solution model for the liquid, face-centered cubic (fcc), body-centered cubic (bcc), and close-packed hexagonal (cph) phases. The  $C14$  phase ( $\text{Fe}_2\text{Ti}$ ) was described by a three sub-lattice model to account for its wide homogeneity range. The slight solubility of Ti in  $\text{Fe}_3\text{P}$  and  $\text{Fe}_2\text{P}$  and the homogeneity ranges of  $\text{Ti}_5\text{P}_3$  and  $\text{TiP}$  were taken into account, by using a two sublattice model. The other binary compounds were modeled as stoichiometric phases. The small homogeneity range of the ternary compound  $\text{FeTiP}$  (anti  $\text{PbCl}_2$ -type orthorhombic) was accounted for, by adopting a three sublattice model. The enthalpies of formation of the binary phosphides and the ternary phosphide  $\text{FeTiP}$  were evaluated from the electron band energy calculations, using the Full Potential Linearized Augmented Plane Wave (FLAPW) method. See [2006Oht] for details.

The experimental phase diagram data of [1959Vog] and [1965Kan] were used in the optimization and the derived interaction parameters were listed. Four vertical sections along the  $\text{Fe}_3\text{P}$ - $\text{FeTiP}$  and  $\text{Fe}$ - $\text{FeTiP}$  joins and at constant 7.5 mass % P and 7.5 mass % Ti respectively, computed by [2006Oht] are compared with the experimental data of [1959Vog] in Fig. 1-4. The agreement is fair. A liquidus

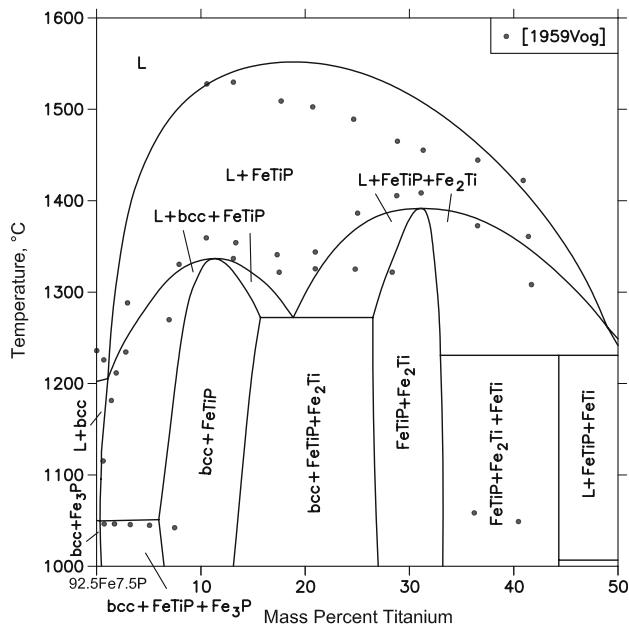


**Fig. 2** Fe-P-Ti computed vertical section along the  $\text{Fe}$ - $\text{FeTiP}$  join [2006Oht]

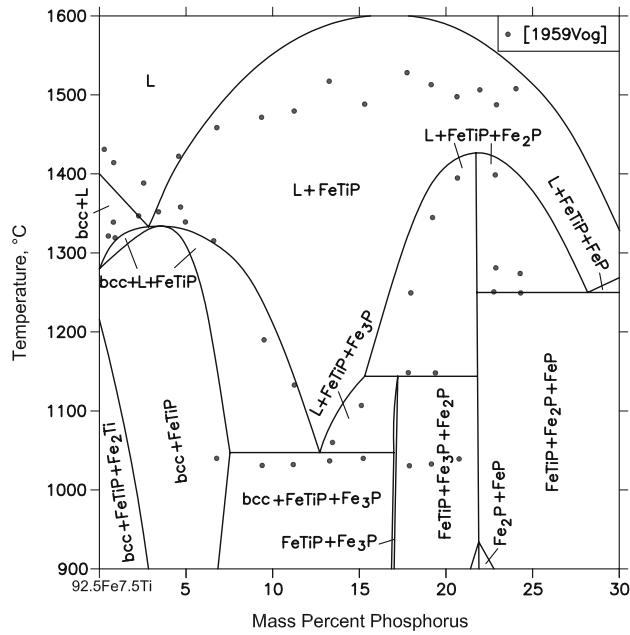
## Section II: Phase Diagram Evaluations

projection for the entire composition range was computed. The computed part in the Fe-FeP-Ti region agrees with the assessed projection of [1988Rag]. Full isothermal sections were also computed at 800 and 25 °C and compared with

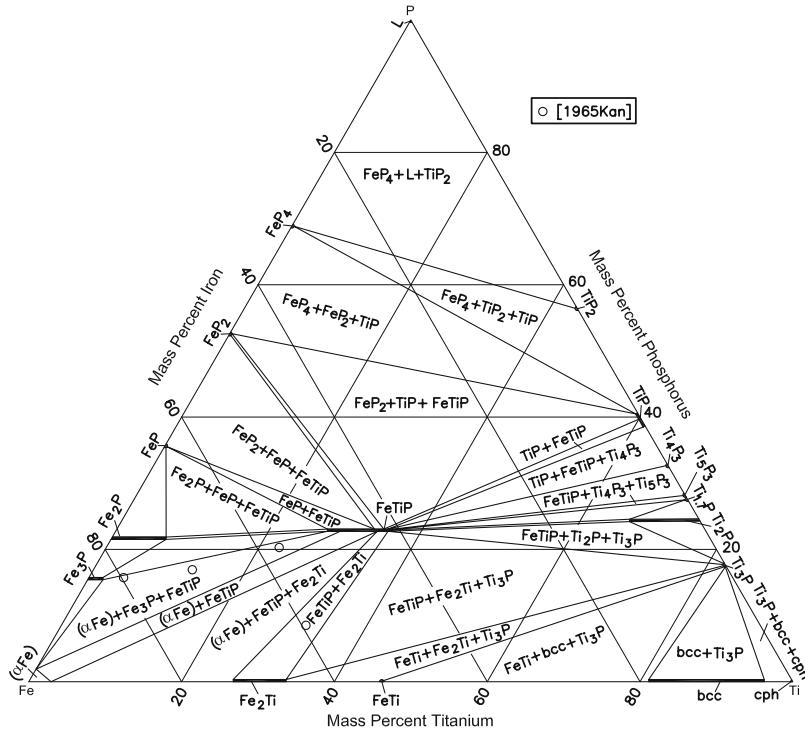
the partial sections at 800 °C by [1965Kan] and 25 °C by [1959Vog]. In Fig. 5, the full computed section at 800 °C [2007Oht] is compared with the few experimental points of [1965Kan].



**Fig. 3** Fe-P-Ti computed vertical section at 7.5 mass % P [2006Oht]



**Fig. 4** Fe-P-Ti computed vertical section at 7.5 mass % Ti [2006Oht]



**Fig. 5** Fe-P-Ti computed isothermal section at 800 °C [2007Oht]

## References

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